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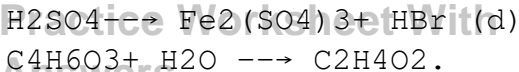
Practice Problems. Balancing  
Equations: Practice

Problems. 1. Balance each of  
the following equations. (a)

$\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3$ . (b)  $\text{Fe} +$

$\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ . (c)  $\text{FeBr}_3 +$

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#2. 1. Potassium reacts with

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water yielding potassium  
hydroxide and hydrogen. 2.

Chlorine reacts with  
potassium bromide yielding  
potassium chloride and  
bromine. 3. Zinc + hydrogen  
chloride yields zinc  
chloride and hydrogen. 4.

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iron + water  $\rightarrow$  Fe<sub>3</sub>O<sub>4</sub> +  
hydrogen  
Answers

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- PRACTICE SHEET # 2

Solutions for the Balancing  
Equations Practice

Worksheet. 1)  $2 \text{NaNO}_3 + \text{PbO}$

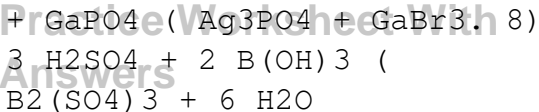
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Answers

1)  $\text{Pb}(\text{NO}_3)_2 + \text{Na}_2\text{O} \rightarrow 2 \text{NaNO}_3 + \text{PbO}$   
2)  $6 \text{AgI} + \text{Fe}_2(\text{CO}_3)_3 \rightarrow 2 \text{FeI}_3 + 3 \text{Ag}_2\text{CO}_3$   
3)  $\text{C}_2\text{H}_4\text{O}_2 + 2 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$   
4)  $\text{ZnSO}_4 + \text{Li}_2\text{CO}_3 \rightarrow \text{ZnCO}_3 + \text{Li}_2\text{SO}_4$   
5)  $\text{V}_2\text{O}_5 + 5 \text{CaS} \rightarrow 5 \text{CaO} + \text{V}_2\text{S}_5$   
6)  $\text{Mn}(\text{NO}_2)_2 + \text{BeCl}_2 \rightarrow \text{Be}(\text{NO}_2)_2 + \text{MnCl}_2$   
7)  $3 \text{AgBr}$

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chemical reaction using

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chemical equations, it is



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important that you become  
firmly acquainted with the  
various part of one. Every  
balanced chemical equation  
consists of two parts: the  
reactant side and the  
product side. Both of these  
sides are separated by the

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means of an arrow. On the  
left side of the arrow, you  
will find the reactant side.

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Balance a chemical equation.

STO.2 Identify the parts of

a chemical equation. RXN.1

Describe a chemical reaction

using words and symbolic

equations. For each of the

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following problems, write complete chemical equations to describe the chemical process taking place. Balance the equations.

Balancing Equations  
Worksheet - 3-13

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Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow. This is a collection of printable worksheets to practice

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your preference below and  
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 $3 \text{Al}_2\text{O}_3 + \text{Fe} \rightarrow \text{Fe}_3\text{O}_4 + \text{Al} \dots$

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