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following questions

in the space

provided. 1. Given

the following

equation: $C_3H_4(g) +$

$xO_2(g) \rightarrow 3CO_2(g) +$

$2H_2O(g)$ 4 a. What is

the value of the

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coefficient x in this equation? 40.07 g/mol b. What is the molar mass of $C_3H_4O_2$? 2 mol O_2 :1 mol H_2O c. What is the mole ratio of O_2 to H_2

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Microscopic: Two

molecules of

hydrogen peroxide

(in aqueous solution)

decompose to

produce two

molecules of liquid

water and one

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SECTION 3 PROBLEMS

Write the answer on
the line to the left.

Show all your work in
the space provided.

1. _____ The actual
yield of a reaction is
22 g and the
theoretical yield is 25
g. Calculate the

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percentage yield. 2.

6.0 mol of N_2 are mixed with 12.0 mol of H

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Theoretically, how many moles of NH_3 will be produced?

PROBLEMS Write the answer on the line to the left, Show all your

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work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: ...

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example is included
below: $2\text{H}_2\text{O}_2(\text{aq})$
 $2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$

This describes the
decomposition
reaction of hydrogen
peroxide.

Microscopic: Two
molecules of
hydrogen peroxide
(in aqueous solution)
decompose to
produce two
molecules of liquid

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water and one molecule of oxygen gas.

Answers

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ANSWER Answer the

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in the space

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the following

equation: $C_3H_4(g) +$

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$x\text{O}_2(\text{g}) + 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar

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ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a.

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