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$I H P_{vap} = \exp 14.790 \cdot T K$ and by the ideal gas law f Solutions to Chemical and Engineering
Thermodynamics, 3e $N V R T \times 18 \text{ mol} \times 8.314 \times 10^5 \text{ (bar} \cdot \text{m}^3 \text{ mol}^{-1} \text{K}^{-1})$ $T P = = V V 1 \times 10^3 \cdot (10 \cdot x) 106$
volume taken up by liquid Equating P and P_{vap} we have $x 18 \times 8.314 \times 10^5 T F 5432.8 \times I$.

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be found if we combine Equations 1.4 and 1.5: $1 \cdot 2 \text{ mV ekmolecular} \cdot 2 \cdot 3 \text{ kT ekmolecular} \cdot 2 \cdot 3 \text{ kT V m}$. Assume
the temperature is 22 C.

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